

Title	A Reversible Rocksalt to Amorphous Phase Transition Involving Anion Redox
Author(s)	Sakuda, Atsushi; Ohara, Koji; Kawaguchi, Tomoya; Fukuda, Katsutoshi; Nakanishi, Koji; Arai, Hajime; Uchimoto, Yoshiharu; Ohta, Toshiaki; Matsubara, Eiichiro; Ogumi, Zempachi; Kuratani, Kentaro; Kobayashi, Hironori; Shikano, Masahiro; Takeuchi, Tomonari; Sakaebe, Hikari
Citation	Scientific Reports (2018), 8
Issue Date	2018-10-10
URL	http://hdl.handle.net/2433/241727
Right	© The Author(s) 2018. This article is licensed under a Creative Commons Attribution 4.0 International License, which permits use, sharing, adaptation, distribution and reproduction in any medium or format, as long as you give appropriate credit to the original author(s) and the source, provide a link to the Creative Commons license, and indicate if changes were made. The images or other third party material in this article are included in the article 's Creative Commons license, unless indicated otherwise in a credit line to the material. If material is not included in the article 's Creative Commons license and your intended use is not permitted by statutory regulation or exceeds the permitted use, you will need to obtain permission directly from the copyright holder. To view a copy of this license, visit http://creativecommons.org/licenses/by/4.0/ .
Type	Journal Article
Textversion	publisher

SCIENTIFIC REPORTS

OPEN

A Reversible Rocksalt to Amorphous Phase Transition Involving Anion Redox

Atsushi Sakuda^{1,6}, Koji Ohara^{2,3}, Tomoya Kawaguchi^{2,7}, Katsutoshi Fukuda², Koji Nakanishi^{2,9}, Hajime Arai^{2,8}, Yoshiharu Uchimoto⁴, Toshiaki Ohta⁵, Eiichiro Matsubara², Zempachi Ogumi², Kentaro Kuratani¹, Hironori Kobayashi¹, Masahiro Shikano¹, Tomonari Takeuchi¹ & Hikari Sakaebe¹

The charge-discharge capacity of lithium secondary batteries is dependent on how many lithium ions can be reversibly extracted from (charge) and inserted into (discharge) the electrode active materials. In contrast, large structural changes during charging/discharging are unavoidable for electrode materials with large capacities, and thus there is great demand for developing materials with reversible structures. Herein, we demonstrate a reversible rocksalt to amorphous phase transition involving anion redox in a Li_2TiS_3 electrode active material with NaCl-type structure. We revealed that the lithium extraction during charging involves a change in site of the sulfur atom and the formation of S—S disulfide bonds, leading to a decrease in the crystallinity. Our results show great promise for the development of long-life lithium insertion/extraction materials, because the structural change clarified here is somewhat similar to that of optical phase-change materials used in DVD-RW discs, which exhibit excellent reversibility of the transition between crystalline and amorphous phase.

As next generation batteries with high energy densities, lithium/sulfur batteries have been widely studied¹. Among these are lithium/metal-polysulfide batteries, which have high theoretical energy densities and utilize metal polysulfides instead of sulfur or lithium sulfide^{2–9}. Conventional lithium/sulfur batteries suffer from a few drawbacks, such as a low volumetric energy density, because they require a large amount of carbon to add electronic conductivity. The choice of electrolytes for lithium/sulfur batteries is limited, owing to adverse side-reactions. In contrast to sulfur electrodes, metal polysulfides have several advantages, such as a high energy density, because of their higher specific gravity and electronic conductivity. Furthermore, a large range of potential electrolytes can also be used in lithium/metal polysulfide batteries, insofar as metal polysulfides show low solubility in many organic solvents, and because undesirable side-reactions are suppressed⁷.

Mechanochemically-prepared Li_2TiS_3 and Li_3NbS_4 have a cation disordered cubic rocksalt structure and can charge/discharge with a capacity greater than 400 mAh g^{−1}⁵. Upon 3.0 V (vs. Li^+/Li) charging and 1.5 V discharging, approximately 2.6 and 3.5 lithium ions can reversibly insert into and extract from the structure with a composition range of $0.4 < x < 3.0$ for Li_xTiS_3 and $0.4 < y < 4.0$ for Li_yNbS_4 , respectively. These materials show excellent cycle performance in all-solid-state lithium cells^{5,6}. Amorphization occurs during the extraction of lithium from Li_2TiS_3 , whereas reverse crystallization occurs during the lithium insertion process. Interestingly, the cation

¹Research Institute of Electrochemical Energy, Department of Energy and Environment, National Institute of Advanced Industrial Science and Technology (AIST), 1-8-31 Midorigaoka, Ikeda, Osaka, 563-8577, Japan. ²Office of Society-Academia Collaboration for Innovation, Kyoto University, Gokasho, Uji, Kyoto, 611-0011, Japan. ³The Research & Utilization Division, Japan Synchrotron Radiation Research Institute (JASRI), 1-1-1 Kouto, Sayo, Hyogo, 679-5198, Japan. ⁴Graduate School of Human and Environmental Studies, Kyoto University, Nihonmatsu-cho, Yoshida, Sakyo-ku, Kyoto, 606-8317, Japan. ⁵SR Center, Ritsumeikan University, 1-1-1 Noji-Higashi, Kusatsu, Shiga, 525-8577, Japan. ⁶Present address: Department of Applied Chemistry, Graduate School of Engineering, Osaka Prefecture University, 1-1 Gakuen-cho, Naka-ku, Sakai, Osaka, 599-8531, Japan. ⁷Present address: Institute for Materials Research, Tohoku University, 2-1-1 Katahira, Aoba-ku, Sendai, Miyagi, 980-8577, Japan. ⁸Present address: Department of Chemical Science and Engineering, School of Materials and Chemical Technology, Tokyo Institute of Technology, 4259 Nagatsuta, Midori-ku, Yokohama, 226-8502, Japan. ⁹Present address: SR Center, Ritsumeikan University, 1-1-1 Noji-Higashi, Kusatsu, Shiga, 525-8577, Japan. Correspondence and requests for materials should be addressed to A.S. (email: saku@chem.osakafu-u.ac.jp) or T.T. (email: takeuchi.tomonari@aist.go.jp)

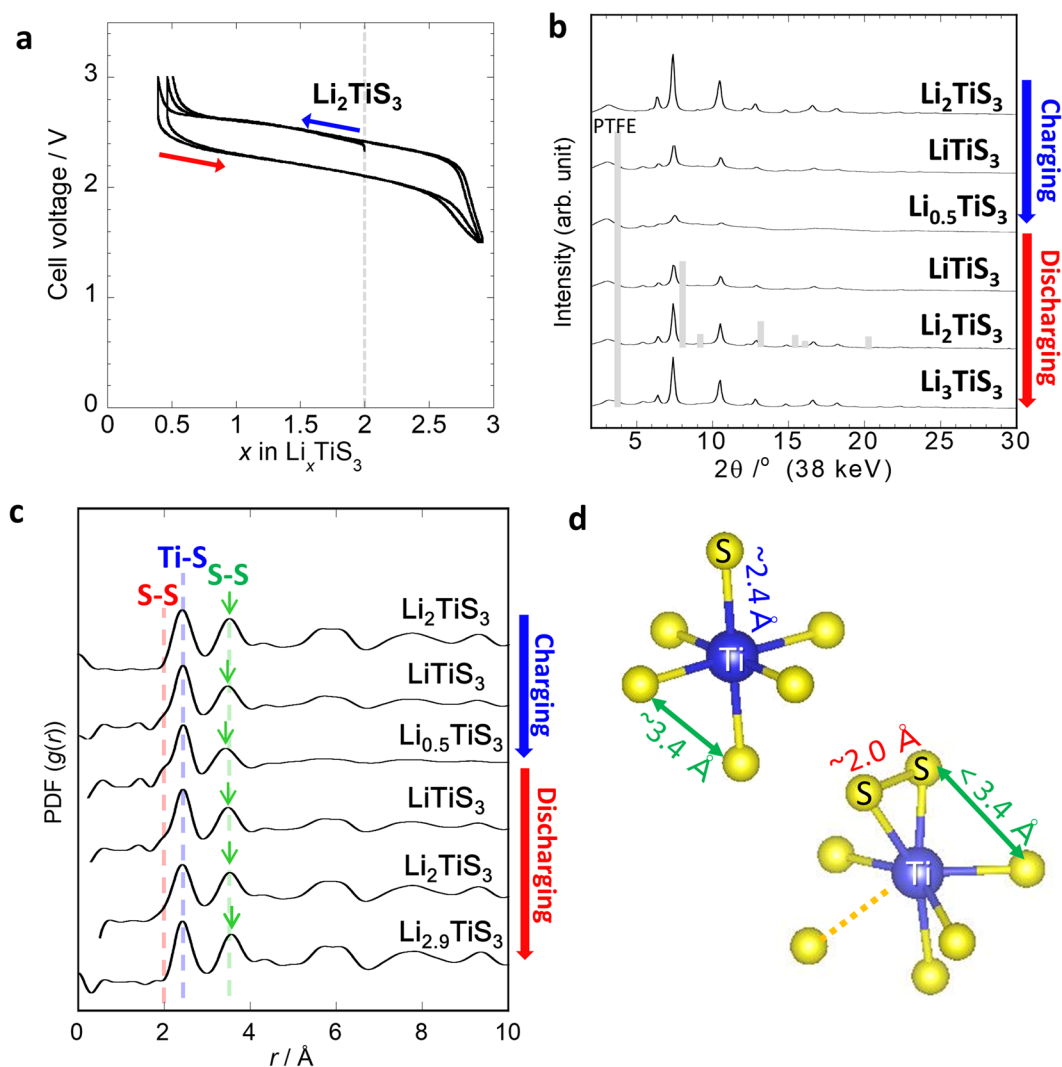


Figure 1. (a) Charge-discharge curves of cubic Li_2TiS_3 , (b) XRD patterns after charging and discharging of cubic Li_2TiS_3 , (c) PDF profiles after charging and discharging of cubic Li_2TiS_3 , and (d) local structure model of Li_2TiS_3 before (upper) and after (bottom) charging.

disordered cubic rocksalt-type phase, which should be metastable at room temperature, is re-formed after lithium insertion. Moreover, the phase transition should accompany anion redox. The nature of the rocksalt/amorphous phase transition has not yet been clarified. Elucidating the mechanisms of the charge-discharge process and structural change offers promise for designing long-life electrode materials with high energy densities because the elucidation of a novel charge-discharge mechanism and its applications have been a major breakthrough in their design, leading to new ideas for boosting their performance^{7,10–16}. We previously reported the reversible anion redox in the amorphous titanium polysulfide electrode by X-ray pair distribution function (PDF) analyses and ab initio molecular dynamics (AIMD) calculations⁷. X-ray PDF analysis is a powerful technique for investigating the structure of amorphous materials. Thus, it is useful to elucidate the structural changes involving amorphization.

Here, we reveal a mechanism of the transition between the rocksalt and amorphous phases in Li_2TiS_3 , including noticeable anion redox by PDF analyses and AIMD calculations.

Results and Discussion

Figure 1 shows the charge-discharge curves, the X-ray diffraction (XRD) patterns, and the PDF profiles following charging and discharging of cubic Li_2TiS_3 . The pattern shows that Li and Ti randomly share the cation site, as we reported previously⁵. The peak intensities for cubic Li_2TiS_3 decreased when charging and increased when discharging, and a shift in the peak positions was barely observed, showing that the Li_2TiS_3 became amorphous upon charging, which agrees with our previous report⁵. The PDF analysis showed similar results; the middle-range order (>5 Å) decreased during charging and increased during discharging. The analysis of the correlations of atoms at short range (<5 Å) reveals the local structural changes involved in the amorphization. Based on an AIMD simulation, we previously reported that the S–S distance in homopolar S–S bonds was 2.0 Å, the Ti–S distance was around 2.4 Å, and the S–S distance of S coordinating to the same Ti was around 3.4 Å, as illustrated in Fig. 1d⁷. These assignments can also be estimated from the crystal structure and Shannon's radii and basically

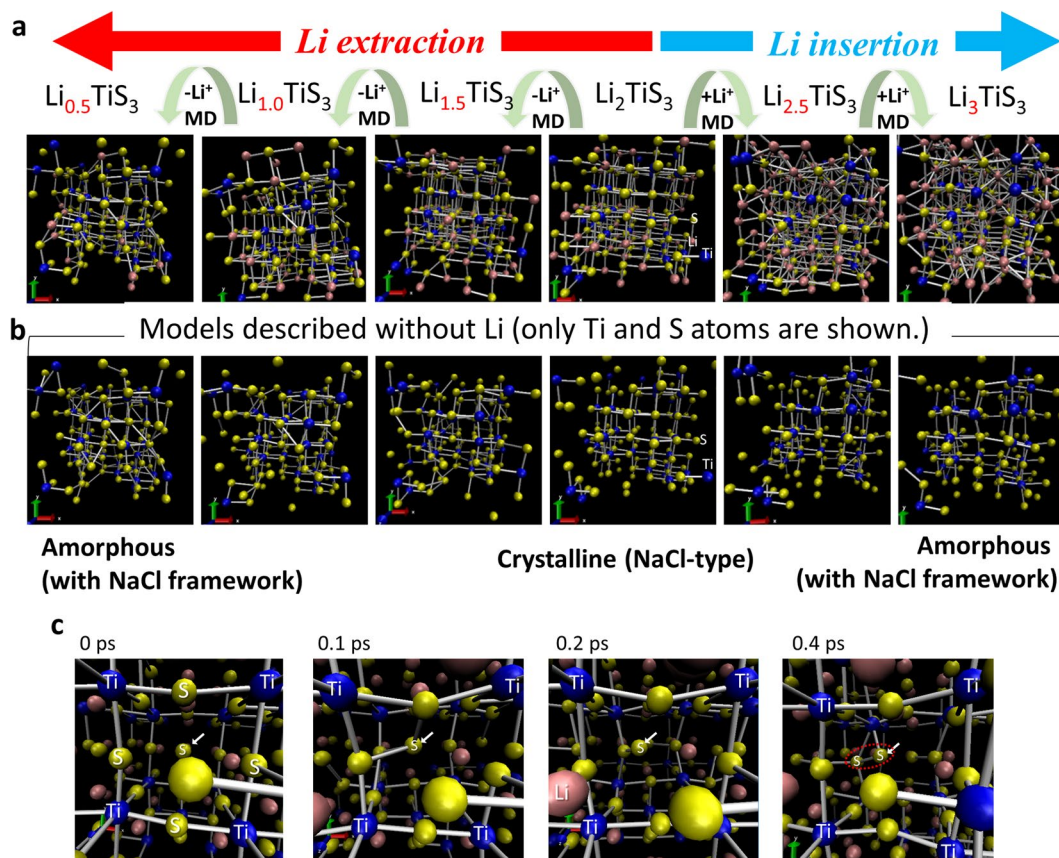


Figure 2. (a) Structural model of Li_xTiS_3 prepared by lithium extraction or insertion using AIMD simulations; (b) models described without Li (only Ti and S atoms are shown); and (c) snapshots of the shifted sulfur during MD simulations.

agreed with the results of AIMD simulations in this study. The peak (shoulder) of the S—S bonds appeared upon charging and disappeared upon discharging, indicating that the covalent S—S bonds reversibly formed and dissociated during the charge-discharge process. These analyses clearly showed two important results: (i) the charge-discharge mechanism involves a phase transition between the disordered rocksalt and amorphous phases, with the short- to middle-range structural change being essentially reversible in a wide composition range from $\text{Li}_{0.5}\text{TiS}_3$ to Li_3TiS_3 , and (ii) the S—S bonds are formed and dissociated during charging and discharging, respectively, showing that the charge-discharge mechanism involves the redox of sulfur. The PDF analyses showed another unusual result: the Ti—S distance remains essentially unchanged during charging and discharging; however, the S—S distance at around 3.4 Å slightly shortens and lengthens during charging and discharging, respectively. This suggests that the coordination environments of both Ti for S and S for Ti change only slightly during charging and discharging. That is, the coordination number of Ti for S was kept at nearly 6 after charging. This is in contrast to the behavior of amorphous TiS_4 electrodes. Indeed, we previously reported that the coordination number of S around Ti continually changes in an amorphous TiS_4 electrode⁷. The X-ray absorption near edge structure (XANES) spectra of S K-edge of Li_2TiS_3 and the charge-discharge products were shown in Fig. S1. The reversible changes in the peak intensity and a shift of the absorption edge energy were observed in the spectra, indicating that the electron state of sulfur reversibly changed during charging and discharging.

We employed computational methods to visualize structural changes during charging/discharging, and these methods are effective at showing the changes that occur in low-crystalline materials⁷. In this study, the AIMD simulations offered a suitable model to explain the experimental results. Figure 2 shows the AIMD simulations of the structural model of Li_xTiS_3 prepared by lithium extraction/insertion. We constructed a $3 \times 3 \times 3$ super cell of cubic rocksalt Li_2TiS_3 with 216 atoms and extracted randomly selected Li atoms to evaluate the structure of $\text{Li}_{2-x}\text{TiS}_3$. After cell parameter optimization, the structure was relaxed by AIMD simulations to obtain the structural models. The lithium excess models were calculated similarly. Figure 2a,b show that the atom ordering in the rocksalt structure decreased with lithium extraction. That is, crystallinity decreased, with the structure becoming amorphous. Conversely, the positions of Ti, Li, and other S atoms were nearly unchanged and the nature of the bonding between them was also almost unchanged. Some of the S atoms moved and formed S—S bonds upon lithium extraction, as shown in Fig. 2c. These results clearly explained why only a slight shift in the XRD peaks was observed and the Ti—S distance was almost unchanged in the PDF profiles; the framework (rocksalt-like structure) formed by the Ti—S—Ti network did not change dramatically. The randomness of Li/Ti and the maintained framework also contributed to the amorphization. The disorder of Li/Ti induced a variation

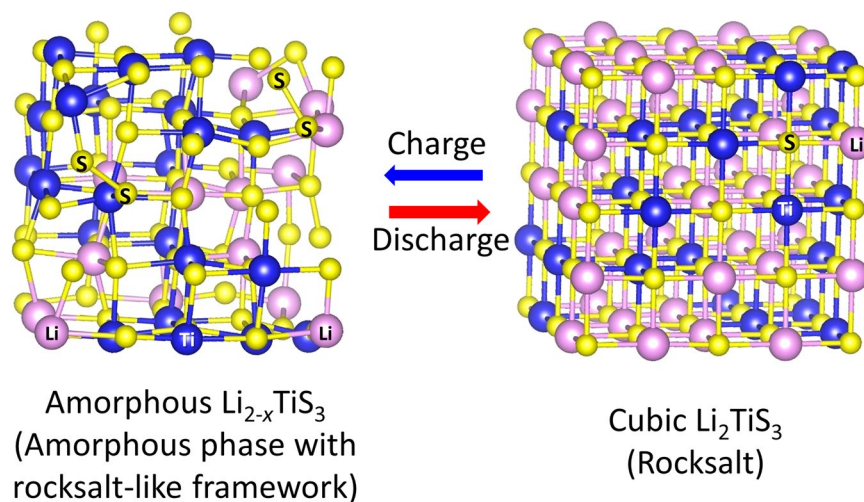


Figure 3. Schematic of structural change during charging and discharging.

in the Li potential, resulting in random extraction of Li with a lower electrochemical potential from the rocksalt structure during charging. As Li atoms were randomly extracted, the rearrangement of the structure was suppressed, and thus its amorphous structure was maintained. This is why the phase transition occurred between the crystalline and amorphous phases rather than the crystalline and crystalline phases. These results indicated that the lithium-extracted Li_xTiS_3 was the material between the crystal and amorphous phases, and that the crystallinity changed gradually and continuously with the Li composition. The schematics of the structural models of amorphous $\text{Li}_{2-x}\text{TiS}_3$ and cubic Li_2TiS_3 are shown in Fig. 3. Electrode materials with a large capacity unavoidably undergo large and usually irreversible structural changes during lithium insertion/extraction, due to the large change in volume. In the phase transition shown here, the volume change is somewhat suppressed as a result of increasing the free volume.

As expected, S atoms that lost more than one neighbour Li preferentially changed their positions to form homopolar S—S bonds, which means that S is oxidized during charging. Furthermore, most disulfides (S—S) directly interact with Ti or Li. Before and after relaxation, the partial density of states (PDOS) of S atoms which form S—S bonds dramatically changed (see Fig. S2), indicating that the charge was compensated for by the S atoms that formed S—S bonds (see Figs S2 and S3, and Table S1). Furthermore, the change in partial atomic charge (Mulliken charge) also showed that the loss of charge due to lithium extraction was compensated for by all the atoms surrounding the extracted lithium. In the cation disordered lithium transition metal sulfides with a lithium-rich composition, the tetrahedron site becomes a diffusion pathway when next neighbour of the site is surrounded by Li, and when the volume of the tetrahedron site is large, as reported by Lee *et al.*¹⁷. Amorphization increases the free volume and distorts the tetrahedral site, facilitating the extraction of Li. This is the reason why a large amount of Li can be extracted from the cubic rocksalt structure.

Unlike in the case of many oxide materials, disulfides formed in the sulfide are accommodated in the framework, and thus showed high reversibility even in deep charging. The results obtained by computational studies were consistent with the experimental results of XRD and PDF analyses. The mechanism elucidated here will be applicable to the study of oxygen¹⁸ and other anion redox couples, depending on the conditions.

With the Li-rich composition of $\text{Li}_{2+x}\text{TiS}_3$, the number of cations becomes larger than that of anions. In this case, it is assumed that a rocksalt structure can no longer be maintained, and that a large structural phase transition will occur. Nevertheless, the rocksalt phase was detected in this composition. It is thought that the peaks in the XRD pattern in Fig. 1b can largely be attributed to Li_2TiS_3 and amorphous $\text{Li}_{2+x}\text{TiS}_3$ formed in the discharged electrodes, rather than crystalline $\text{Li}_{2+x}\text{TiS}_3$. The AIMD simulation suggests a decrease in crystallinity, as shown in Fig. 2. Further studies are needed to determine whether the phase containing additional lithium is amorphous or whether sulfur defects are involved.

Interestingly, the transition between the rocksalt and amorphous phases examined in this study partially resembles that in optical phase transition materials, which are used in DVD-RW and DVD-RAM discs^{19–21}. In a representative optical phase transition material—viz., $\text{Ge}_2\text{Sb}_2\text{Te}_5$ —the Ge—Te—Ge network is a core network in both crystalline (disordered rocksalt) and amorphous phases, which contributes to rapid and highly reversible phase transition. The Sb—Te bond is relatively weak and fragile, and acts as a trigger for crystallization and amorphization²¹. The role of the Ti—S—Ti network in this study is similar to that of the Ge—Te—Ge core network in the phase transition material. The behaviour of the Li—S and S—S bonds is also similar to that of the Sb—Te bond. In optical phase change materials, rapid and excellent reversibility has been achieved over 10^6 cycles. Thus, it is expected that high-power electrode materials with an ultra-long cycle life can be developed based on the idea of mimicking optical phase-change materials.

Conclusion

In this paper, we described the mechanism of the reversible rocksalt to amorphous phase transition involving anion redox in cubic Li_2TiS_3 . Extracting lithium while charging involves the oxidation of sulfur and the formation of S–S disulfide bonds; the formation of a disulphide bond leads to a decrease in crystallinity. The reverse structural change occurs when discharging. Although cubic Li_2TiS_3 is a metastable phase, the phase is preferably formed during lithium insertion after amorphization by lithium extraction, because the amorphous phase contains a cubic rocksalt analogue Ti–S–Ti core framework. The core framework forms an intermediate state between amorphous and crystal. We also found similarities between the phase transition in the electrode material of this study and that of optical phase transition materials. Anion redox combined with the mimicry of optical phase-change materials is thus a promising approach to the future development of novel large-capacity and ultra-long-life electrode materials.

Methods

Materials. Cubic Li_2TiS_3 was prepared by a mechanochemical process⁵. Crystalline TiS_2 (99.8%, Aldrich) and Li_2S (99.9%, Mitsuwa Pure Chemicals) were used as the starting materials. Zirconia pots (45 mL) and 500 zirconia balls (4 mm in diameter) were used for ball-milling; the planetary ball-mill apparatus (P-7, Fritsch) was operated at 510 rpm for 40 h.

Electrochemical cells. To prepare the Li_2TiS_3 working electrode, Li_2TiS_3 , acetylene black (AB) and polytetrafluoroethylene (PTFE) were mixed in an agate mortar⁵. The weight ratio of Li_2TiS_3 , AB and PTFE was 9:1:1. The working electrode was attached to an Al mesh current corrector. The loading rate of the active material was $\sim 20 \text{ mg cm}^{-2}$. A 1 M solution of LiPF_6 in a 50:50 (by volume) mixture of ethylene carbonate and dimethyl carbonate (Tomiya Pure Chemical Industries Ltd.) was used as the electrolyte, and a piece of Li foil was used as the counter electrode. The electrochemical measurements were performed at 30 °C using a charge-discharge unit (TOSCAT-3100, Toyo System) at a current density of 20 mA g^{-1} .

High-energy X-ray diffraction measurements. The X-ray total scattering measurements for the PDF analysis were carried out at Spring-8 BL28XU²². The incident X-ray energy was 38.0 keV, and a Si (220) crystal ($2d = 3.840 \text{ \AA}$) monochromator was used. The analysed Q-range was $0.7\text{--}14 \text{ \AA}^{-1}$. The intensity of the incident X-rays was monitored in an ionisation chamber filled with Ar gas, while the scattered X-rays were detected with a YAP detector. A vacuum chamber was used to suppress air scattering. The collected datasets were corrected for the absorption, background and polarisation effects. Details of the correction and normalisation procedures have been reported elsewhere²³.

DFT-MD simulations. The AIMD were performed using OpenMX (ver. 3.8)²⁴. The exchange-correlation functional used was GGA-PBE. The cut-off energy was 300 Ry. The basis functions (cut-off radius/orbital structure) used were 7.0 Bohr/ $s^3p^3d^2$ (Ti), 7.0 Bohr/ $s^2p^2d^1$ (S) and 8.0 Bohr/ s^3p^2 (Li). A norm-conserving pseudopotential was used. For the MD calculations, the O(N) Krylov-subspace method was used²⁵. The MD temperature was controlled to 400 K, and the time step was 1.2 fs/MD-step. The FOCUS supercomputer system was used for the calculations. The atomic models were described using VESTA software²⁶.

$3 \times 3 \times 3$ super cells of cubic rock-salt Li_2TiS_3 (216 atoms) were constructed by total energy calculation. The lowest energy configuration in the calculated model was determined from 20 different configurations. Randomly selected Li atoms were then extracted and the cell parameter with the lowest energy was recalculated. Finally, the structure was relaxed by AIMD at 400 K and optimized at 0 K. The lithium extraction and relaxation process was repeated several times.

X-ray absorption fine structure (XAFS) measurements. The sulfur K-edge XAFS measurements were carried out at BL-10 of the SR Center, Ritsumeikan University. The partial fluorescence yield (PFY) mode was used to obtain the spectra with a silicon drift detector. The incident X-ray beam was monochromatised with a Ge(111) crystal ($2d = 6.532 \text{ \AA}$). The absolute photon energy was calibrated with the assumption that the strong resonance of K_2SO_4 (S 1s $\rightarrow t_2$) appears at 2481.7 eV.

References

- Ji, X. & Nazar, L. F. Advances in Li–S batteries. *J. Mater. Chem.* **20**, 9821–9826 (2010).
- Hayashi, A., Matsuyama, T., Sakuda, A. & Tatsumisago, M. Amorphous Titanium Sulfide Electrode for All-solid-state Rechargeable Lithium Batteries with High Capacity. *Chem. Lett.* **41**, 886–888 (2012).
- Matsuyama, T. *et al.* Preparation of amorphous TiS_x thin film electrodes by the PLD method and their application to all-solid-state lithium secondary batteries. *J. Mater. Sci.* **47**, 6601–6606 (2012).
- Sakuda, A. *et al.* Amorphous TiS_4 positive electrode for lithium-sulfur secondary batteries. *Electrochem. Commun.* **31**, 71–75 (2013).
- Sakuda, A. *et al.* Rock-salt-type lithium metal sulphides as novel positive-electrode materials. *Sci. Rep.* **4**, 4883 (2014).
- Sakuda, A., Takeuchi, T., Shikano, M., Sakaebe, H. & Kobayashi, H. High Reversibility of ‘Soft’ Electrode Materials in All-Solid-State Batteries. *Front. Energy Res.* **4**(19), 1–7 (2016).
- Sakuda, A. *et al.* Amorphous Metal Polysulfides: Electrode Materials with Unique Insertion/Extraction Reactions. *J. Am. Chem. Soc.* **139**, 8796–8799 (2017).
- Doan-Nguyen, V. V. T. *et al.* Molybdenum polysulfide chalcogenides as high-capacity, anion-redox-driven electrode materials for Li-ion batteries. *Chem. Mater.* **28**, 8357–8365 (2016).
- Wang, X. *et al.* Unique Reversible Conversion-Type Mechanism Enhanced Cathode Performance in Amorphous Molybdenum Polysulfide. *ACS Appl. Mater. Interfaces* **9**, 38606–38611 (2017).
- Poizot, P., Laruelle, S., Grugeon, S., Dupont, L. & Tarascon, J. M. Nano-sized transition-metal oxides as negative-electrode materials for lithium-ion batteries. *Nature* **407**, 496–499 (2000).
- Tarascon, J. M. & Armand, M. Issues and challenges facing rechargeable lithium batteries. *Nature* **414**, 359–367 (2001).
- Armand, M. & Tarascon, J. M. Building better batteries. *Nature* **451**, 652–657 (2008).

13. Dunn, B., Kamath, H. & Tarascon, J.-M. Electrical Energy Storage for the Grid: A Battery of Choices. *Science* **334**, 928–935 (2011).
14. Aricò, A. S., Bruce, P., Scrosati, B., Tarascon, J. & van Schalkwijk, W. Nanostructured materials for advanced energy conversion and storage devices. *Nat. Mater.* **4**, 366–377 (2005).
15. Jung, S.-K. *et al.* Lithium-free transition metal monoxides for positive electrodes in lithium-ion batteries. *Nat. Energy* **2**, 16208 (2017).
16. Sakuda, A. *et al.* Cubic Rocksalt Li_2SnS_3 and a Solid Solution with Li_3NbS_4 Prepared by Mechanochemical Synthesis. *Electrochemistry* **85**, 580–584 (2017).
17. Lee, J. *et al.* Unlocking the Potential of Cation-Disordered Oxides for Rechargeable Lithium Batteries. *Science* **343**, 519–522 (2014).
18. Yabuuchi, N. *et al.* Origin of stabilization and destabilization in solid-state redox reaction of oxide ions for lithium-ion batteries. *Nat. Commun.* **7**, 13814 (2016).
19. Oguchi, H. *et al.* Experimental and computational studies on structural transitions in the LiBH_4 -LiI pseudobinary system. *Appl. Phys. Lett.* **94**, 141912 (2009).
20. Wuttig, M. & Yamada, N. Phase-change materials for rewriteable data storage. *Nat. Mater.* **6**, 824–832 (2007).
21. Ohara, K. *et al.* The roles of the Ge-Te core network and the Sb-Te pseudo network during rapid nucleation-dominated crystallization of amorphous $\text{Ge}_2\text{Sb}_2\text{Te}_5$. *Adv. Funct. Mater.* **22**, 2251–2257 (2012).
22. Tanida, H. *et al.* RISING beamline (BL28XU) for rechargeable battery analysis. *J. Synchrotron Radiat.* **21**, 268–272 (2014).
23. Kohara, S. *et al.* Structural studies of disordered materials using high-energy x-ray diffraction from ambient to extreme conditions. *J. Phys. Condens. Matter* **19**, 506101 (2007).
24. Ozaki, T. *et al.* OpenMX. <http://www.openmx-square.org/> (2016).
25. Ozaki, T. O(N) Krylov-subspace method for large-scale ab initio electronic structure calculations. *Phys. Rev. B* **74**, 245101–245117 (2006).
26. Momma, K. & Izumi, F. VESTA 3 for three-dimensional visualization of crystal, volumetric and morphology data. *J. Appl. Cryst.* **44**, 1272–1276 (2011).

Acknowledgements

This work was supported by the RISING and RISINGII projects of the New Energy and Industrial Technology Development Organization (NEDO), Japan.

Author Contributions

A.S., T.T. and H.S. conceived the study; A.S. synthesised the materials and wrote the manuscript; A.S., K.O., K.F. and K.N. carried out the experiments; Z.O., E.M., Y.U., T.O., H.S. and H.K. supervised the project; all authors analysed the data and commented on the manuscript.

Additional Information

Supplementary information accompanies this paper at <https://doi.org/10.1038/s41598-018-33518-4>.

Competing Interests: The authors declare no competing interests.

Publisher's note: Springer Nature remains neutral with regard to jurisdictional claims in published maps and institutional affiliations.



Open Access This article is licensed under a Creative Commons Attribution 4.0 International License, which permits use, sharing, adaptation, distribution and reproduction in any medium or format, as long as you give appropriate credit to the original author(s) and the source, provide a link to the Creative Commons license, and indicate if changes were made. The images or other third party material in this article are included in the article's Creative Commons license, unless indicated otherwise in a credit line to the material. If material is not included in the article's Creative Commons license and your intended use is not permitted by statutory regulation or exceeds the permitted use, you will need to obtain permission directly from the copyright holder. To view a copy of this license, visit <http://creativecommons.org/licenses/by/4.0/>.

© The Author(s) 2018